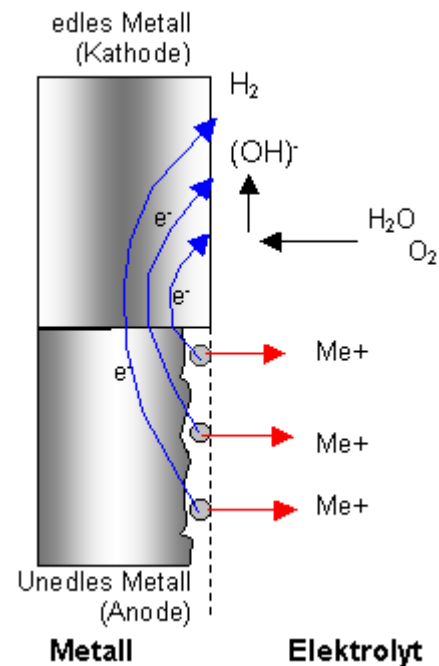
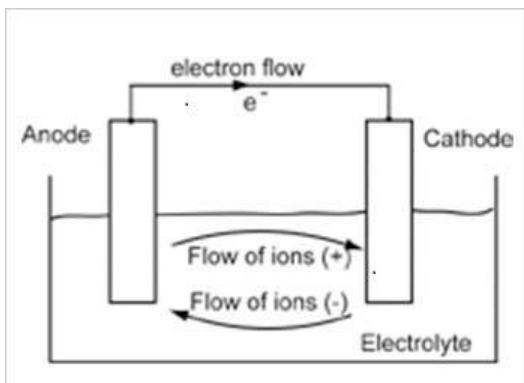


Galvanic Corrosion

Galvanic corrosion occurs when different, interconnected metals are surrounded by a common electrolyte (sea water, water, humidity). The metal with the higher electrical potential becomes the anode, and the one with the lower becomes the cathode. Current flows from the anode to the cathode. The anode will dissolve or corrode.



The position of the metals in the galvanic series is decisive for the corrosion process. The further apart they are in this row, the faster and more aggressively the corrosion progresses.

Galvanic voltage series

- Magnesium
- Zinc
- Aluminum
- Iron
- Lead
- Tinplate
- Chrome alloys
- Brass
- Copper
- Nickel
- Stainless Steel
- Silver
- Gold

